

Terbium

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Terbium is a chemical element with symbol **Tb** and atomic number 65. It is a silvery-white rare earth metal that is malleable, ductile and soft enough to be cut with a knife. The ninth member of the lanthanide series, terbium is a fairly electropositive metal that reacts with water, evolving hydrogen gas. Terbium is never found in nature as a free element, but it is contained in many minerals, including cerite, gadolinite, monazite, xenotime and euxenite.

Swedish chemist Carl Gustaf Mosander discovered terbium as a separate elemental compound in 1843. He detected it as an impurity in yttrium oxide, Y₂O₃. Yttrium and terbium are named after the village of Ytterby in Sweden.

Terbium was not isolated in pure form until the advent of ion exchange techniques.

Terbium is used to dope calcium fluoride, calcium tungstate and strontium molybdate, materials that are used in solid-state devices, and as a crystal stabilizer of fuel cells which operate at elevated temperatures. As a component of Terfenol-D (an alloy that expands and contracts when exposed to magnetic fields more than any other alloy), terbium is of use in actuators, in naval sonar systems and in sensors.

Most of the world's terbium supply is used in green phosphors. Terbium oxide is in fluorescent lamps and TV tubes. Terbium green phosphors are combined with divalent europium blue phosphors and trivalent europium red phosphors to provide "trichromatic" lighting technology, a high-efficiency white light used for standard illumination in indoor lighting.

Characteristics

Physical properties

Terbium, ⁶⁵Tb



General properties

Name, symbol terbium, Tb

Appearance silvery white

Terbium in the periodic table

Atomic number (Z) 65

Group, block group n/a, f-block

Period period 6

Element category ☐ lanthanide

Standard atomic weight (\pm) (*A*_r) 158.92535(2)^[1]

Electron configuration [Xe] 4f⁹ 6s²

per shell 2, 8, 18, 27, 8, 2

Physical properties

Phase solid

Melting point 1629 K (1356 °C, 2473 °F)

Boiling point 3396 K (3123 °C, 5653 °F)

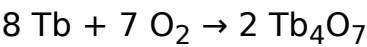
Terbium is a silvery-white rare earth metal that is malleable, ductile and soft enough to be cut with a knife.^[2] It is relatively stable in air compared to the earlier, more reactive lanthanides in the first half of the lanthanide series.^[3] Terbium exists in two crystal allotropes with a transformation temperature of 1289 °C between them.^[2] The 65 electrons of a terbium atom are arranged in the electron configuration [Xe]4f⁹6s²; normally, only three electrons can be removed before the nuclear charge becomes too great to allow further ionization, but in the case of terbium, the stability of the half-filled [Xe]4f⁷ configuration allows further ionization of a fourth electron in the presence of very strong oxidizing agents such as fluorine gas.^[2]

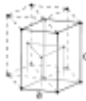
The terbium(III) cation is brilliantly fluorescent, in a bright lemon-yellow color that is the result of a strong green emission line in combination with other lines in the orange and red. The yttrifluorite variety of the mineral fluorite owes its creamy-yellow fluorescence in part to terbium. Terbium easily oxidizes, and is therefore used in its elemental form specifically for research. Single terbium atoms have been isolated by implanting them into fullerene molecules.^[4]

Terbium has a simple ferromagnetic ordering at temperatures below 219 K. Above 219 K, it turns into a helical antiferromagnetic state in which all of the atomic moments in a particular basal plane layer are parallel, and oriented at a fixed angle to the moments of adjacent layers. This unusual antiferromagnetism transforms into a disordered paramagnetic state at 230 K.^[5]

Chemical properties

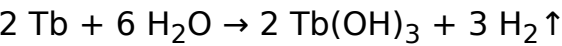
The most common oxidation state of terbium is +3, as in Tb₂O₃. The +4 state is known in TbO₂ and TbF₄.^{[6][7]} Terbium burns readily to form a mixed terbium(III,IV) oxide:^[8]



Density near r.t.	8.23 g/cm ³					
when liquid, at m.p.	7.65 g/cm ³					
Heat of fusion	10.15 kJ/mol					
Heat of vaporization	391 kJ/mol					
Molar heat capacity	28.91 J/(mol·K)					
Vapor pressure						
P (Pa)	1	10	100	1 k	10 k	100 k
at T (K)	1789	1979	(2201)	(2505)	(2913)	(3491)
Atomic properties						
Oxidation states	4, 3 , 2, 1 (a weakly basic oxide)					
Electronegativity	Pauling scale: 1.2 (?)					
Ionization energies	1st: 565.8 kJ/mol 2nd: 1110 kJ/mol 3rd: 2114 kJ/mol					
Atomic radius	empirical: 177 pm					
Covalent radius	194±5 pm					
Miscellanea						
Crystal structure	hexagonal close-packed (hcp)					
						
Speed of sound thin rod	2620 m/s (at 20 °C)					
Thermal expansion	at r.t. α, poly: 10.3 μm/(m·K)					
Thermal conductivity	11.1 W/(m·K)					
Electrical resistivity	α, poly: 1.150 μΩ·m (at r.t.)					
Magnetic ordering	paramagnetic at 300 K					



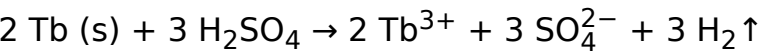
In solution, terbium forms only trivalent ions. Terbium is quite electropositive and reacts slowly with cold water and quite quickly with hot water to form terbium hydroxide:^[8]



Terbium metal reacts with all the halogens, forming white trihalides:^[8]



Terbium dissolves readily in dilute sulfuric acid to form solutions containing the pale pink terbium(III) ions, which exist as a [Tb(OH₂)₉]³⁺ complexes.^[8]



Compounds

Terbium combines with nitrogen, carbon, sulfur, phosphorus, boron, selenium, silicon and arsenic at elevated temperatures, forming various binary compounds such as TbH₂, TbH₃, TbB₂, Tb₂S₃, TbSe, TbTe and TbN.^[7] In those compounds, Tb mostly exhibits the oxidation states +3 and sometimes +2. Terbium(II) halogenides are obtained by annealing Tb(III) halogenides in presence of metallic Tb in tantalum containers. Terbium also forms sesquichloride Tb₂Cl₃, which can

be further reduced to TbCl by annealing at 800 °C. This terbium(I) chloride forms platelets with layered graphite-like structure.^[9]

Other compounds include

- Chlorides: TbCl₃
- Bromides: TbBr₃
- Iodides: Tbl₃
- Fluorides: TbF₃, TbF₄

Terbium(IV) fluoride is a strong fluorinating agent, emitting relatively pure atomic fluorine when heated^[10] rather than the mixture of fluoride vapors emitted from CoF₃ or CeF₄.

Young's modulus

Shear modulus

Bulk modulus

Poisson ratio

Vickers hardness

Brinell hardness

CAS Number

α form: 55.7 GPa

α form: 22.1 GPa

α form: 38.7 GPa

α form: 0.261

450–865 MPa

675–1200 MPa

7440-27-9

History

Naming

Discovery and first isolation

after Ytterby (Sweden), where it was mined

Carl Gustaf Mosander (1842)

Most stable isotopes of terbium

iso	NA	half-life	DM	DE (MeV)	DP
157Tb	syn	71 y	ε	0.060	157Gd
158Tb	syn	180 y	ε	1.220	158Gd
			β−	0.937	158Dy
159Tb	100%	is stable with 94 neutrons			

Isotopes

Naturally occurring terbium is composed of its only stable isotope, terbium-159; the element is thus called mononuclidic and monoisotopic. Thirty six radioisotopes have been characterized, with the heaviest being terbium-171 (with atomic mass of 170.95330(86) u) and lightest being terbium-135 (exact mass unknown).^[11] The most stable synthetic radioisotopes of terbium are terbium-158, with a half-life of 180 years, and terbium-157, with a half-life of 71 years. All of the remaining radioactive isotopes have half-lives that are much less than a quarter of a year, and the majority of these have half-lives that are less than half a minute.^[11] The primary decay mode before the most abundant stable isotope, ¹⁵⁹Tb, is electron capture, which results in production of gadolinium isotopes, and the primary mode after is beta minus decay, resulting in dysprosium isotopes.^[11]

The element also has 27 nuclear isomers, with masses of 141–154, 156, and 158 (not every mass number corresponds to only one isomer). The most stable of them are terbium-156m, with half-life of 24.4 hours and terbium-156m2, with half-life of 22.7 hours; this is longer than half-lives of most ground states of radioactive terbium isotopes, except only those with mass numbers 155–161.^[11]

External links

- Wikipedia: Terbium (<https://en.wikipedia.org/wiki/Terbium>)